COMMUNICATIONS TO THE EDITOR

X-RAY INVESTIGATIONS ON RUBBER

Sir:

During some earlier x-ray investigations on crêpe rubber, it was observed that the originally translucent sample became more and more milky at the spot where the x-rays passed the material. The sample was the same as described in an earlier paper.[1] This rubber gave a distinct x-ray pattern during the first exposure, but, although the lines did not change their position, they became increasingly sharper with each exposure; simultaneously the opaqueness of the milky spot increased. The increase of sharpness of the originally broad lines represents, of course, increase in particle size; but evidently also the increasing cloudiness of the crêpe rubber shows progressing crystallization.

Another sample of the crêpe rubber was placed in front of the window of the x-ray tube and again slowly the milky spot formed becoming eventually entirely opaque toward transmitted light. The spot obtained was sharply defined by the transmitted x-ray beam and showed no diffuseness at the circumference. After four years of storage, no further changes occurred in the sample. Upon slight warming the milky spot would melt away again.

An increase in the rate of crystallization of gutta-percha, but not of rubber, by means of ultraviolet light was observed by Kirchhof.[2]

Evidently the above sample of rubber behaves similarly under the influence of x-rays as does gutta-percha with ultraviolet light.

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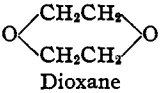
THE RAMAN SPECTRUM OF DIOXANE

Sir:

During the course of an investigation of the Raman spectrum of some synthetic resins, the solvent dioxane was used. As the spectrum of the latter has not been published previously, I wish to report here my findings.



Dioxane

An apparatus designed by Reynolds and Benford [*Review of Scientific Instruments*, **1**, 413–416 (1930)] was used, modified by having a chromium elliptical reflector instead of a silver one, and its major axis in a horizontal rather than a vertical plane. A mercury arc gave the exciting spectrum, exposures of about six hours being made.

[1] E. Ott, *Naturwissenschaften*, **14**, 320 (1926).

[2] Kirchhof, *Kautschuk*, **4**, 254 (1928); *C. A.*, **23**, 1525 (1929).

The reciprocals of the wave lengths of the modified lines were as follows (the intensities are given in parentheses) and are accurate probably to within 6 cm.$^{-1}$.

24426 (0), 225 (0), 23878 (4), 691 (3), 590 (1), 496 (1), 408 (2), 267 (2), 22507 (00), 422 (0), 099 (4), 21997 (00), 930 (1), 853 (3), 744 (5), 21640 (1), 556 (0), 501 (1), 20214 (00), 088 (4), 19972 (5). The lines were correlated, respectively, as follows, the ordinal numbers in the preceding list being given first, followed by the respective wave number shift.

Exciting line								Average $\Delta\tilde{\nu}$
	24710		24522		22941		23047	
1	284 (0)	2	297 (0)	9?	434 (00)			291 (0) cm.$^{-1}$
3	838 (4)	4	831 (3)	10?	519 (0)			837 (4)
5	1120 (1)	7	1114 (2)	11	842 (4)	13	1117 (1)	1117 (1)
6	1214 (1)							1214 (1)
8	1443 (2)			18	1440 (1)			1442 (2)
12?	2713 (00)			19	2727 (00)			2720 (00)
14	2866 (3)	16	2882 (1)	20	2853 (4)			2864 (3)
15	2966 (5)	17	2966 (0)	21	2969 (5)			2967 (3)
								434 (00)?
								519 (0)?

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THE PHOTOCHEMICAL POLYMERIZATION OF ACETYLENE

Sir:

It has been reported by several observers[1] that the polymerization of acetylene to a solid resembling cuprene is induced by either activated mercury atoms or by the direct absorption of ultraviolet light. A quantitative investigation of the photochemical reaction has been undertaken in the hope of throwing some light on the kinetics of the process and to determine whether there is any simple relation between the *quantum yield* and *the ion pair yield* of the analogous radiochemical reaction.[2]

The following are some of the more important results of this investigation. Using the complete radiation of a quartz mercury arc and a reaction vessel three centimeters deep, the reaction velocity was observed to be proportional to the pressure of acetylene as long as this pressure was not greater than 30 cm.; at higher pressures the velocity approached a constant value. It seems probable therefore that the rate of polymerization

[1] Berthelot and Gaudechon, *Compt. rend.*, **150**, 1169 (1910); Bates and Taylor, THIS JOURNAL, **49**, 2437 (1927); Reinike, *Z. angew. Chem.*, **41**, 1144 (1928).

[2] Lind, Bardwell and Perry, THIS JOURNAL, **48**, 1556–1575 (1926); Mund and Koch, *Bull. soc. chim. Belg.*, **34**, 241–255 (1925).